

**GATITU SECONDARY SCHOOL, P.O. BOX 327 – 01030, GATUNDU.**  
**FORM 1 CHEMISTRY TUNE UP EXAMINATION. TERM 3 2015**

**NAME:** \_\_\_\_\_ **CLASS:** \_\_\_\_\_ **ADM. NO.** \_\_\_\_\_

**INSTRUCTION:**

- 1) time 45 Minutes
2. Write down your name, admission Number and Class in the spaces provided
3. Answer all questions in the spaces below each question.
4. 30 Marks

1. Name the elements present in the following compounds.

a) Sodium bromide (1mk)

b) Zinc Sulphate (1 ½ mks)

c) Magnesium Nitride (1mk)

d) Copper (II) Nitrate (1 ½ mks)

2. List three differences between temporary and permanent changes.

TEMPORARY

PERMANENT

(3MKS)

3. Write word equations for the reaction between dilute hydrochloric acid and each one of the following.

a) Zinc metal (2mks)

b) Calcium hydrogen carbonate (2mks)

c) Potassium hydroxide (2mks)

4. Dilute sulphuric acid was added to a compound of magnesium P. The solid reacted with the acid to form a colourless solution Q, and a colourless gas R which formed a white precipitate when bubbled through Calcium hydroxide solution.

a) Name

i) Compound P (1mk)

ii) Solution Q (1mk)

iii) Colourless gas R (1mk)

b) Write a word equation for the reaction that took place. (2mks)

c) State the observations that would be made if a similar compound of calcium was used instead of magnesium. Explain (3mks)

5. Sodium chloride (common salt) is contaminated with copper (II) oxide. Explain how pure sodium chloride can be obtained from the mixture. (3mks)

b) Solution may be classified as strong base, weak base, neutral strong acid or weak acid. The information below gives some solutions and their  $p^H$  values. Study it and answer the questions that follow.

SOLUTION	$p^H$
A	0.5
B	7
C	14
D	9

